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Chapter 12 Stoichiometry127 SECTION 12.1 THE ARITHMETIC OF EQUATIONS (pages 353–358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

### **SECTION 12.1 THE ARITHMETIC OF EQUATIONS**

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## **Chapter 12 - Study Guide - Answers**

Chapter 3 Stoichiometry 3-3 3.1a Avogadro's Number The mole (abbreviated mol) is the unit chemists use when counting numbers of atoms or molecules in a sample. The number of particles (atoms, molecules, or other objects) in one mole is equal to the number of atoms in exactly 12 g of carbon-12.

## **Chapter 3 Stoichiometry**

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Wednesday, February 22 Agenda: 1) Review stoichiometry 2) quiz - stoichiometry HW: 1) write pre lab - Mass of Magnesium in your lab notebook. pdf FREE PDF DOWNLOAD Chapter 12 Stoichiometry Practice Problems Answer Key €. Stoichiometry Worksheet #1 Answers - MR Wu0026#39;s GHHS Science Website Stoichiometry Worksheet #1 Answers 1.

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